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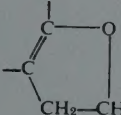
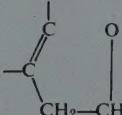
ERRATA

(The line numbers do not take account of titles and headings)

Page 30, column 2, line 18	read <i>Epitrix</i>	for <i>Epitix</i>
" 71, " 2, footnote 4	" acetate	" arsenate
" 80, " 1, line 17	" <i>oxysporum</i>	" <i>oxyporum</i>
" 92, name of author	" Gonçalves	" Gonçalves

ERRATA AND ADDITIONS TO BE INCLUDED IN ISSUES OF THE BULLETIN PRIOR TO 1964

(The line numbers do not take account of titles and headings)

Volume	9, page	19, column	2, Table I	read 3(i), 3(ii), 3(iii), 3(iv)	for 3(i), 3(ii), 3(iii), 3(iv)
"	9, "	27, "	1, between lines 20 and 21 insert	Heat under reflux until the mixture is boiling gently, add the condenser zinc wool (1 gm) rolled into a ball, continue boiling for 15 minutes, repeat with a further two lots of zinc wool (3 × 1 gm in all). Remove the source of heat, rinse the condenser with a little acetic acid, allow the apparatus to cool until the flask can be conveniently handled, and detach the flask (Note 2). Fit the Buchner funnel with two 5.5-cm filter papers and cover them with a layer of Gooch asbestos (Notes 1 and 3). Decant the hot contents of the flask on to the filter, taking care to retain the zinc, which may have partly decomposed, in the flask as far as possible. Rinse the flask with water and pass the rinsings through the filter. Wash the residue in the flask and on the filter with water until the washings are halide-free, i.e., give no precipitate on testing with silver nitrate solution.	
"	9, "	154, "	1, line 38	read Determine the chloride ion either by the Volhard method using the thiocyanate or electrometrically.	Titrate the solution with the thiocyanate or determine the chloride ion electrometrically.
"	10, "	67, "	1, structural formula		
"	10, "	112, "	1, line 34	$w/w = \frac{304.45 \times t}{w}$	$w/w = \frac{304.45 - t}{w \times 100}$
"	10, "	112, "	2, " 40	"	"
"	10, "	113, "	1, " 46	"	"
"	10, "	114, "	1, " 8	"	"
"	10, "	134, "	1, " 2	290.9	296.5
"	10, "	135, "	2, " 8	Determine the chloride ion by the Volhard method using the thiocyanate solution or electrometrically.	Titrate the solution with the thiocyanate or determine the chloride ion electrometrically.

Volume 11, page 110, column 1, structural formula

read NO₂for N₂O

" 11, " 111, " 1, after line 10 insert

Determine the melting point of the isolated DNOC by the prescribed method MT 2 (below). This should not be less than 84°C and should not be depressed by admixture with an equal quantity of pure DNOC.

" 11, " 112, " 1, line 2

read 255.5

" 176.9

" 11, " 132, " 2, " 10

" (27 mg)

" (27 gm)